

Jérôme Canivet, Bruno Therrien
 and Georg Süss-Fink*

Institut de Chimie, Université de Neuchâtel,
 Case postale 2, CH-2007 Neuchâtel,
 Switzerland

Correspondence e-mail:
 georg.suess-fink@unine.ch

Key indicators

Single-crystal X-ray study
 T = 173 K
 Mean σ (C–C) = 0.006 Å
 R factor = 0.023
 wR factor = 0.079
 Data-to-parameter ratio = 17.1

For details of how these key indicators were
 automatically derived from the article, see
<http://journals.iucr.org/e>.

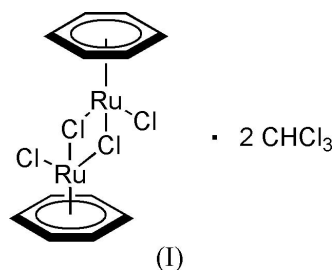
Di- μ -chloro-bis[(η^6 -benzene)chlororuthenium(II)]
 chloroform disolvate

In the centrosymmetric dinuclear title complex, $[\text{RuCl}_2(\eta^6\text{-C}_6\text{H}_6)]_2 \cdot 2\text{CHCl}_3$ or $[\text{Ru}_2\text{Cl}_4(\text{C}_6\text{H}_6)_2] \cdot 2\text{CHCl}_3$, accessible from $\text{RuCl}_3 \cdot n\text{H}_2\text{O}$ and 1,3-cyclohexadiene, the benzene ligands are involved in slipped-parallel π - π stacking interactions with neighbouring molecules, thus forming one-dimensional polymeric chains.

Received 28 April 2005
 Accepted 3 May 2005
 Online 14 May 2005

Comment

The title compound, (I), which has been known for more than 25 years (Bennett & Smith, 1974) and extensively used as a building block in organometallic chemistry, has never been characterized by X-ray structure analysis. It was crystallized here as the chloroform disolvate.



The two halves of the dinuclear complex (Fig. 1) are related by a crystallographic inversion center. The bond distances and angles (Table 1) are similar to those of other $[\text{RuCl}_2(\eta^6\text{-arene})]_2$ complexes; arene = hexamethylbenzene (McCormick & Gleason, 1988), trindane (Gupta *et al.*, 1997), ethylbenzoate (Therrien *et al.*, 1998), 1,2,3,4-tetrahydronaphthalene (Bown & Bennett, 1999), 1,2- $\text{C}_6\text{H}_4(\text{Me})\text{COOMe}$ (Braga *et al.*, 2001), hexaethylbenzene (Baldwin *et al.*, 2002), indane (Vieille-Petit *et al.*, 2002) and *para*-cymene (Allardyce *et al.*, 2003).

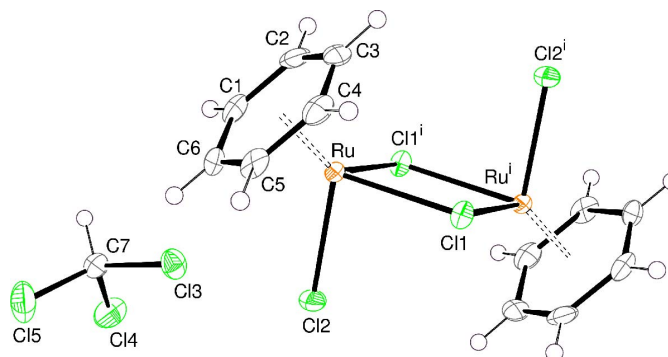


Figure 1
 The molecular structure of (I), showing 50% displacement ellipsoids (arbitrary spheres for the H atoms). The symmetry code is as in Table 1. The centrosymmetrically related chloroform molecule is not shown.

In (I), the benzene ring is planar and the Ru–(benzene centroid) distance is 1.646 Å. The metal possesses two bridging and one terminal chloride ions: the bridging Ru–Cl distances are 2.4495 (11) and 2.4580 (9) Å, and the terminal Ru–Cl distance is 2.3911 (9) Å. The benzene ligands are involved in slipped-parallel π – π stacking interactions with neighbouring molecules (Fig. 2) and the shortest intermolecular C...C distance is 3.218 (5) Å. The distance observed between the centers of the interacting π systems (4.07 Å) is in good agreement with the theoretical value calculated for this π stacking mode (Tsuzuki *et al.*, 2002). No meaningful interactions between the dinuclear ruthenium complex and the chloroform molecule are observed.

Experimental

[RuCl₂(η^6 -C₆H₆)₂]₂ was dissolved in hot chloroform and orange crystals of (I) suitable for X-ray diffraction analysis were obtained after 3 d by slow evaporation of the solvent.

Crystal data

| | |
|---|---|
| [Ru ₂ Cl ₄ (C ₆ H ₆) ₂] ₂ ·2CHCl ₃ | Z = 1 |
| <i>M_r</i> = 738.90 | <i>D_x</i> = 2.219 Mg m ⁻³ |
| Triclinic, <i>P</i> $\bar{1}$ | Mo <i>K</i> α radiation |
| <i>a</i> = 7.9951 (10) Å | Cell parameters from 5113 reflections |
| <i>b</i> = 8.1835 (11) Å | θ = 2.4–25.9° |
| <i>c</i> = 9.6545 (13) Å | μ = 2.57 mm ⁻¹ |
| α = 69.312 (15)° | <i>T</i> = 173 (2) K |
| β = 69.676 (15)° | Block, orange |
| γ = 86.164 (16)° | 0.35 × 0.23 × 0.12 mm |
| <i>V</i> = 553.03 (15) Å ³ | |

Data collection

| | |
|--|--|
| Stoe IPDS diffractometer | 2014 independent reflections |
| φ scans | 1817 reflections with <i>I</i> > 2 σ (<i>I</i>) |
| Absorption correction: refined from ΔF (DIFABS; Walker & Stuart, 1983) | <i>R</i> _{int} = 0.032 |
| <i>T</i> _{min} = 0.397, <i>T</i> _{max} = 0.794 | θ _{max} = 25.9° |
| 4350 measured reflections | <i>h</i> = –9 → 9 |
| | <i>k</i> = –10 → 10 |
| | <i>l</i> = –11 → 11 |

Refinement

| | |
|-------------------------------------|--|
| Refinement on <i>F</i> ² | $w = 1/[\sigma^2(F_o^2) + (0.0458P)^2 + 0.223P]$ |
| $R[F^2 > 2\sigma(F^2)] = 0.023$ | where $P = (F_o^2 + 2F_c^2)/3$ |
| $wR(F^2) = 0.079$ | $(\Delta\sigma)_{\max} < 0.001$ |
| <i>S</i> = 1.20 | $\Delta\rho_{\max} = 0.59 \text{ e } \text{Å}^{-3}$ |
| 2014 reflections | $\Delta\rho_{\min} = -0.88 \text{ e } \text{Å}^{-3}$ |
| 118 parameters | |
| H-atom parameters constrained | |

Table 1

Selected geometric parameters (Å, °).

| | | | |
|-------------------------|------------|--------------------------|-------------|
| Ru–Cl1 | 2.4580 (9) | Ru–Cl1 ¹ | 2.4495 (11) |
| Ru–Cl2 | 2.3911 (9) | | |
| Ru ¹ –Cl1–Ru | 98.22 (3) | Cl1 ¹ –Ru–Cl1 | 81.78 (3) |
| Cl2–Ru–Cl1 | 86.43 (3) | | |

Symmetry code: (i) –*x*, 1 – *y*, 1 – *z*.

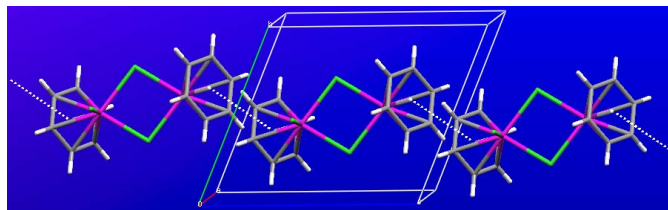


Figure 2

Part of an infinite chain of [RuCl₂(η^6 -C₆H₆)₂]₂ molecules in (I), arising through slipped-parallel π – π stacking interactions.

H atoms were placed in calculated positions (C–H = 0.93 Å) and treated as riding atoms with the constraint $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{carrier})$ applied.

Data collection: *EXPOSE* in *IPDS Software* (Stoe & Cie, 2000); cell refinement: *CELL* in *IPDS Software*; data reduction: *INTEGRATE* in *IPDS Software*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 1990); program(s) used to refine structure: *SHELXL97* (Sheldrick, 1997); molecular graphics: *ORTEP-3* (Farrugia, 1997); software used to prepare material for publication: *SHELXL97*.

This work was supported by the Swiss National Science Foundation (grant No. 20-61227-00). We thank Professor H. Stoeckli-Evans for helpful discussions and free access to X-ray facilities.

References

- Allardycce, C. S., Dyson, P. J., Ellis, D. J., Salter, P. A. & Scopelliti, R. (2003). *J. Organomet. Chem.* **668**, 35–42.
- Baldwin, R., Bennett, M. A., Hockless, D. C. R., Pertici, P., Verrazzani, A., Uccello Barretta, G., Marchetti, F. & Salvadori, P. (2002). *J. Chem. Soc. Dalton Trans.* pp. 4488–4496.
- Bennett, M. A. & Smith, A. K. (1974). *J. Chem. Soc. Dalton Trans.* pp. 233–240.
- Bown, M. & Bennett, M. A. (1999). *Acta Cryst.* **C55**, 852–854.
- Braga, D., Abati, A., Scaccianoce, L., Johnson, B. F. G. & Grepioni, F. (2001). *Solid State Sci.* **3**, 783–788.
- Farrugia, L. J. (1997). *J. Appl. Cryst.* **30**, 565.
- Gupta, H. K., Lock, P. E., Hughes, D. W. & McGlinchey, M. J. (1997). *Organometallics*, **16**, 4355–4361.
- McCormick, F. B. & Gleason, W. B. (1988). *Acta Cryst.* **C44**, 603–605.
- Sheldrick, G. M. (1990). *Acta Cryst.* **A46**, 467–473.
- Sheldrick, G. M. (1997). *SHELXL97*. University of Göttingen, Germany.
- Stoe & Cie (2000). *IPDS Software*. Stoe & Cie GmbH, Darmstadt, Germany.
- Therrien, B., Ward, T. R., Pilkington, M., Hoffmann, C., Gilardoni, F. & Weber, J. (1998). *Organometallics*, **17**, 330–337.
- Tsuzuki, S., Honda, K., Uchimura, T., Mikami, M. & Tanabe, K. (2002). *J. Am. Chem. Soc.* **124**, 104–112.
- Vieille-Petit, L., Therrien, B. & Süß-Fink, G. (2002). *Acta Cryst.* **E58**, m656–m657.
- Walker, N. & Stuart, D. (1983). *Acta Cryst.* **A39**, 158–166.